Methanol-Ethanol Mixture

Methanol and ethanol have vapor pressures of 159 and 58 Torr, respectively at 298 K. Determine the pressure at which the last drop of solution will go into the vapor phase for a 0.5 mole fraction mixture of the two solvents.

Methanol-Ethanol Mixture

Methanol and ethanol have vapor pressures of 159 and 58 Torr, respectively at 298 K. Determine the pressure at which the last drop of solution will go into the vapor phase for a 0.5 mole fraction mixture of the two solvents.

Solution: The last drop will go into the vapor phase when the pressure has reached the curve calculated using Dalton's law.

$$P_{tot} = \frac{P_1^* P_2^*}{P_2^* - y_2 (P_2^* - P_1^*)}$$

$$P_{tot} = \frac{(58)(159)}{159 - 0.5(159 - 58)}$$

$$P_{tot} = 85 Torr$$