Enthalpy of combustion

Compare the enthalpy of combustion of H_2, N_2 and Cl_2. You will need the following data.

$$\Delta_f H^o(NO_2) = +33.6 \ kJ/mol$$
$$\Delta_f H^o(ClO_2) = +104.8 \ kJ/mol$$
$$\Delta_f H^o(H_2O) = -286 \ kJ/mol$$

Of the diatomic molecules given here explain why is H_2 a good fuel.