## Constructing a reaction table



Construct a reaction table for a mixture of 1.9 grams of $\mathrm{CaCl}_{2}$ with 1.9 grams of $\mathrm{Na}_{3} \mathrm{PO}_{4}$. Assuming that the equilibrium constant is extremely large for this reaction determine the limiting reagent and then determine the mass of each reactant remaining and each product produced after the reaction has taken place. The unbalanced chemical reaction is:

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-\mathrm{CaCl}_{2}+\_\mathrm{Na}_{3} \mathrm{PO}_{4} \rightleftarrows \_\mathrm{Ca}_{3}\left(\mathrm{PO}_{4}\right)_{2}+\_\mathrm{NaCl}
$$

