By Henry's law we know that $\mathrm{CO}_{2}$ will dissolve in $\mathrm{H}_{2} \mathrm{O}$ at a mole fraction of $\mathrm{x}_{\mathrm{CO}_{2}}=1.2 \times 10^{-5}$. Calculate the molarity of dissolved $\mathrm{CO}_{2}$ under these conditions.

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Solution: This solution is very dilute. Therefore, we are in the limit $x_{1} \sim 1$ for the solvent. Therefore,

$$
c_{2}=x_{2}(55.56 M)
$$

which yields

$$
c_{2}=\left(1.2 \times 10^{-5}\right)(55.56 M)=6.67 \times 10^{-4} M
$$

