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Solution: In this type of problem we can calculate the number of moles of solvent in 1 L, which we already know is 55.6 moles for H_2O. Then we can calculate the number of moles of $CuSO_4$

$$\mathrm{n_2 = c_2(1\ L) = (4.3\ M)(1\ L) = 4.3\ mol}$$

Thus, we can calculate

$$x_{CuSO_4} = \frac{n_{CuSO_4}}{n_{CuSO_4} + n_{H_2O}} = \frac{4.3}{4.3 + 55.6} = 0.071$$