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Solution: The molality is measured by kg of solvent. Therefore, we can use the fact that a 25% solution Contains 333 gm of anti-freeze. We need to convert this value into moles and then we are done since that will be the number of moles per 1000 g.

$$n_{C_2O_2H_6} = \frac{m_{C_2O_2H_6}}{M_{m,C_2O_2H_6}} = \frac{333\ gm}{62\ gm/mol} = 5.37\ mol$$

The solution is 5.37 molal in ethylene glycol.